Physical Science | Atomic Box and Atomic Structure | Date

Ms. Toal

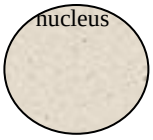
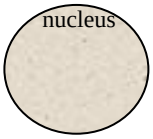
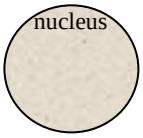
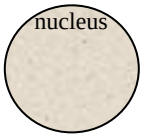
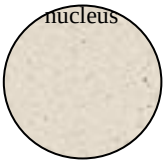
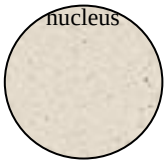
Kenilworth JHS | Name | Period

A. p+ n e- DRAWINGS

1. For each box fill in the blanks
2. Fill in the nucleus with the correct number of particles(use the key)
3. Draw in the correct number of shells
4. Draw in the correct number and placement of electrons on the shells

KEY

+= Protons
o= Neutrons
-= Electrons

3 **Li** 6.941 p+=____ e-=____ n=____  Rounded atomic mass ___ - atomic #___= ___n	12 **Mg** 24.305 p+=____ e-=____ n=____  Rounded atomic mass ___ - atomic #___= ___n
9 **F** 18.998 p+=____ e-=____ n=____  Rounded atomic mass ___ - atomic #___= ___n	7 **N** 14.007 p+=____ e-=____ n=____  Rounded atomic mass ___ - atomic #___= ___n
15 **P** 30.974 p+=____ e-=____ n=____  Rounded atomic mass ___ - atomic #___= ___n	18 **Ar** 39.948 p+=____ e-=____ n=____  Rounded atomic mass ___ - atomic #___= ___n

B. You have been assigned element # _____

For this element you must completely fill out the atomic box correctly. Use your notes in your lab book to make sure you do this correctly for your assigned element. After you complete the atomic box, you must illustrate the atomic structure for this same element (draw the correct location and the correct number of protons, neutrons and electrons).

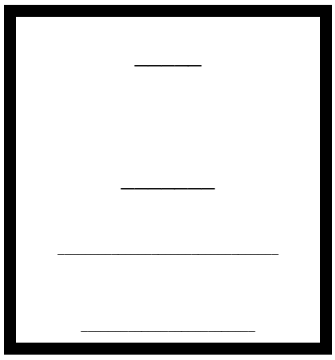
Element name = _____

Element Symbol = _____ Atomic Number = _____

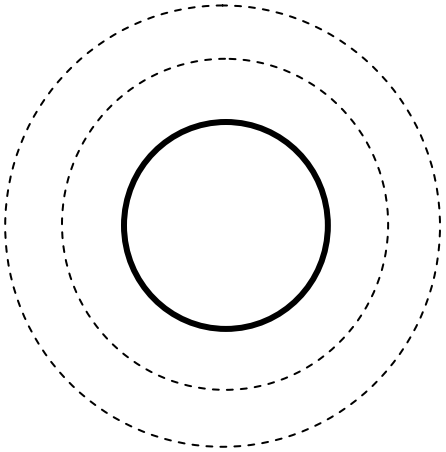
Number of Protons = _____ Atomic Mass = _____

Number of Neutrons = _____

Number of Electrons = _____



In the area below, draw the atomic structure for your element (correct location and number of protons, neutrons, and electrons).



C. Use your knowledge of atomic number and mass number and fill in the missing numbers:

Element	Atomic Number	Atomic Mass	Mass Number (Atomic mass rounded)	Protons	Neutrons	Electrons
Sodium	11		23		12	
Phosphorus	15				16	15
Magnesium	12		24			
Carbon			12	6		
Potassium	19	39.10				
Lithium	3					
Argon	18					
Copper	29					
Mercury		201				
Silver					61	47