

# VALENCE ELECTRONS

Copy what is underlined

# 1. Valence Electrons

- ⦿ A. The electrons that are *farthest* away from the nucleus
- ⦿ B. These e- are special because *they* are involved in *chemical reactions/bonding*

◎ C. Held most *loosely*

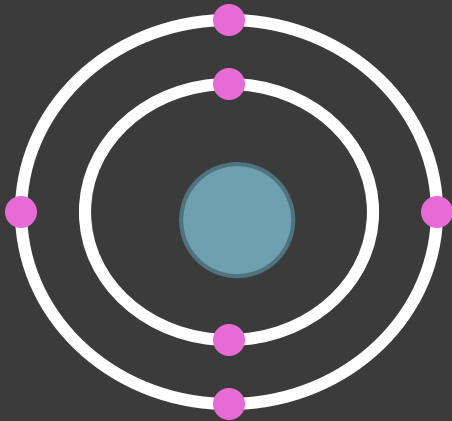
◎ D. Can *transfer*

2. Elements in the same  
group have:

- Same number of valence  
electrons

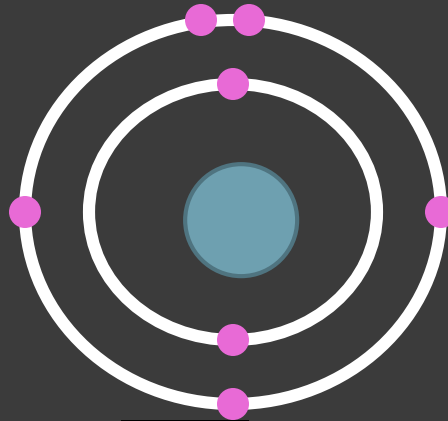
# 3. ELECTRON DOT DIAGRAMS

CARBON



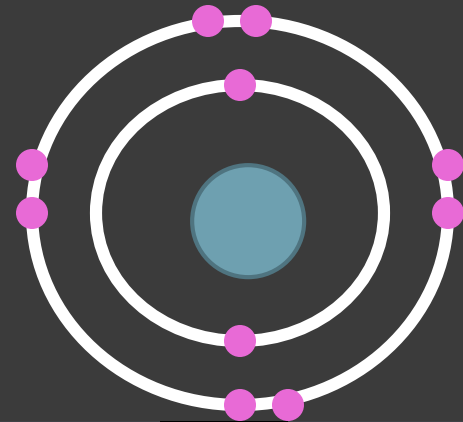
4 Valence e-

NITROGEN



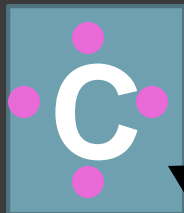
5

NEON



8

Draw  
valence  
e- around  
symbol

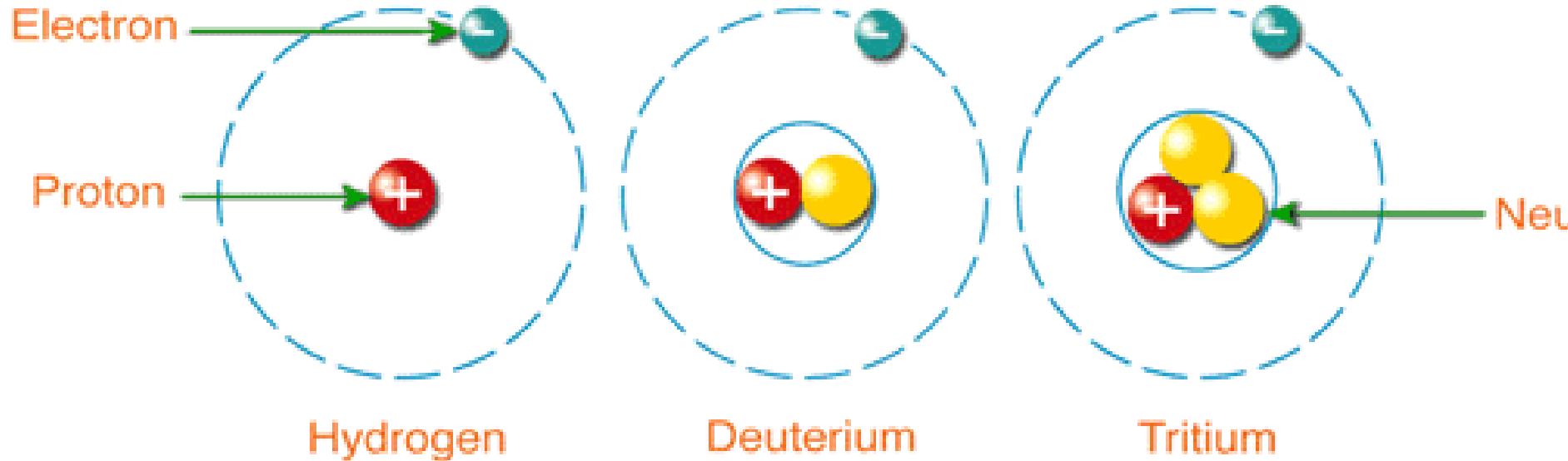


ELECTRON DOT  
DIAGRAM

## 4. Extra info

- ⦿ A. Dark zigzag line on PT goes through semimetals
- ⦿ B. Elements in the same row/period, have the same number of shells
- ⦿ C. Isotopes have the same number of  $p^+$ , but different number of  $n$

# ISOTOPE



- ⦿ What is overall charge of each?
  - NEUTRAL or 0
  - What particle is different?
    - NEUTRONS